

Gold Nanocomposites Based on 1-Vinyl-1,2,4-Triazole Copolymer with Acrylic Acid

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ISSN: 2770-6613



Abstract

The paper discusses the results of electrosynthesis of metal-polymer nanocomposites of gold and their coating of pure iron and steel electrodes when combining the process of electro polymerization of 1-vinyl-1.2.4-triazole with acrylic acid during cathodic separation of metals. Using UV, IR and atomic absorption spectroscopies, X-ray phase analysis, as well as thermogravimetric and elemental analyses, the structure and composition of the synthesized nanocomposites and nanocomposite coatings were studied.

Keywords: Electrolysis; Nanocomposites; 1-vinyl-1.2.4-triazole; Gold; Electro polymerization

Introduction

Electrosynthesis and study of the properties of functional polymers and nanomaterials obtained on their basis is the most rapidly developing area of modern chemical science. As noted in [1-5], nanomaterials are used in medicine as antibacterial drugs, systems for targeted delivery of contrast agents and drugs, biosensors, and other biomedical purposes. Nanocomposite materials containing gold nanoparticles have unique properties and their development is promising for medicine, nanophotonic and catalysis [1,2].

Mini Review

The synthesis of new functional thrombo-resistant, non-toxic polymeric materials with gold nanoparticles makes it possible to expand the area of their application, as well as to increase the range of materials used in pharmaceuticals when creating new dosage forms. The properties of nanocomposites substantially depend on both the nature of the stabilizing polymer matrix and the conditions for the nanoparticle's formation. As a polymer matrix, 1-Vinyl-1.2.4-Triazole (VT) polymers and copolymers can be used, since they are non-toxic ($LD_{50} > 3000 \text{ mg} \cdot \text{kg}^{-1}$) and have high film formation ability, solubility, and biocompatibility [6,7]. In this work, we discuss the results of electrosynthesis of metal-polymer nanocomposites and their coatings on pure iron and steel electrodes when the process of electrochemical (co) polymerization of solid solution with Acrylic Acid (AA) is combined with cathodic separation of metals. During the electrolysis of aqueous or water-ethanol solutions of VT and AA or their mixtures at various ratios in the presence of HAuCl_4 and chitosan, nanocomposites, and nanocomposite coatings with a gold content of 1-10wt% are formed only in the presence of a peroxide-type initiator, for example, TBOBA, the potential of electroreduction of which is close to the potentials of the cathodic separation of metals - 0.6-1.2V (c.s.e.). After drying, the formed nanocomposite coatings on the electrodes become insoluble in water and in commonly used organic solvents (dimethylsulfoxide, dimethylformamide, acetonitrile, etc.). Likely, the copolymer is crosslinked upon heating. In the electronic spectra of gold-containing nanocomposites, in contrast to aqueous solutions of the initial copolymers and HAuCl_4 , plasmon absorption bands appear with a maximum in the region of nanocomposite coatings containing gold, absorption bands appear with a maximum in the region of 517-521 nm, which is typical for systems with zero-valent gold.

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Submission:  March 08, 2022

Published:  March 18, 2022

Volume 3 - Issue 2

How to cite this article: Sargsyan SA*, Sargsyan AS, Sargsyan TS, Khizantsyan KM, Agajanyan IG, et al. Gold Nanocomposites Based on 1-Vinyl-1,2,4-Triazole Copolymer with Acrylic Acid. *Polymer Sci peer Rev J.* 3(2). PSPRJ. 000560. 2022.
DOI: [10.31031/PSPRJ.2022.03.000560](https://doi.org/10.31031/PSPRJ.2022.03.000560)

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The IR spectra of VT-AA copolymers contain bands corresponding to the frequencies of stretching bending vibrations of the triazole ring $1503, 1434, 1138, 1005, 660\text{cm}^{-1}$ (C-N, C=N), 1275cm^{-1} (NN), 3106cm^{-1} (CH), and a band at 1711cm^{-1} related to stretching vibrations of the units of the carboxyl group (Figure 1). Analysis of IR spectra shows that the formation of gold-containing polymer nanocomposites leads to insignificant changes in the chemical structure of the copolymer matrix. Thus, the intensities of the absorption bands of the triazole ring, which can act as coordination centers of gold nanoparticles, show a weak shift (by $3\text{-}4\text{cm}^{-1}$) of one band at 1506cm^{-1} , which is characteristic of the stretching vibrations of the ring (C-N and C=N). This shift may indicate the

coordination interaction of the triazole ring with the surface atoms of metal nanoparticles. As a result of this work, a nanocomposite of the following structure is formed (Figure 2). The sizes of gold nanoparticles were calculated according. According to the results of transmission electron microscopy (Figure 3a), nanocomposites contain mostly elliptical gold nanoparticles, uniformly distributed in the polymer matrix, having sizes of $2\text{-}10\text{nm}$ (Figure 3b). The study of the thermal stability of polymer nanocomposites showed that the first stage of polymer matrix destruction is observed in the temperature range from $260\text{ }^{\circ}\text{C}$ to $400\text{ }^{\circ}\text{C}$ and is accompanied by a gradual weight loss of up to 45% , which, likely, refers to the elimination and oxidation of carboxyl (Figure 4).

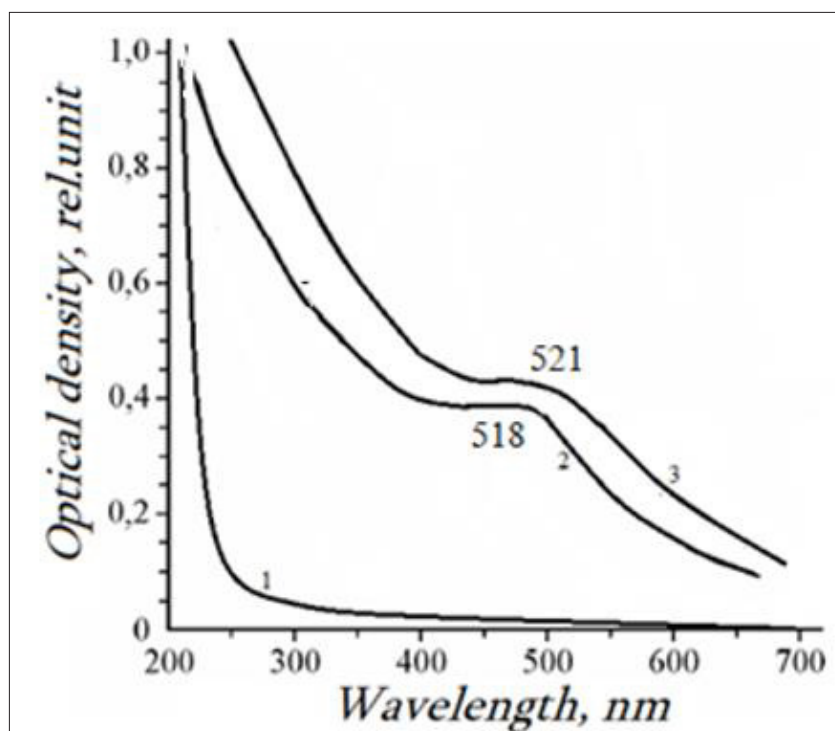


Figure 1: Electronic absorption spectra of VT-AA copolymer (1), 4.8% (2), 6.1% (3).

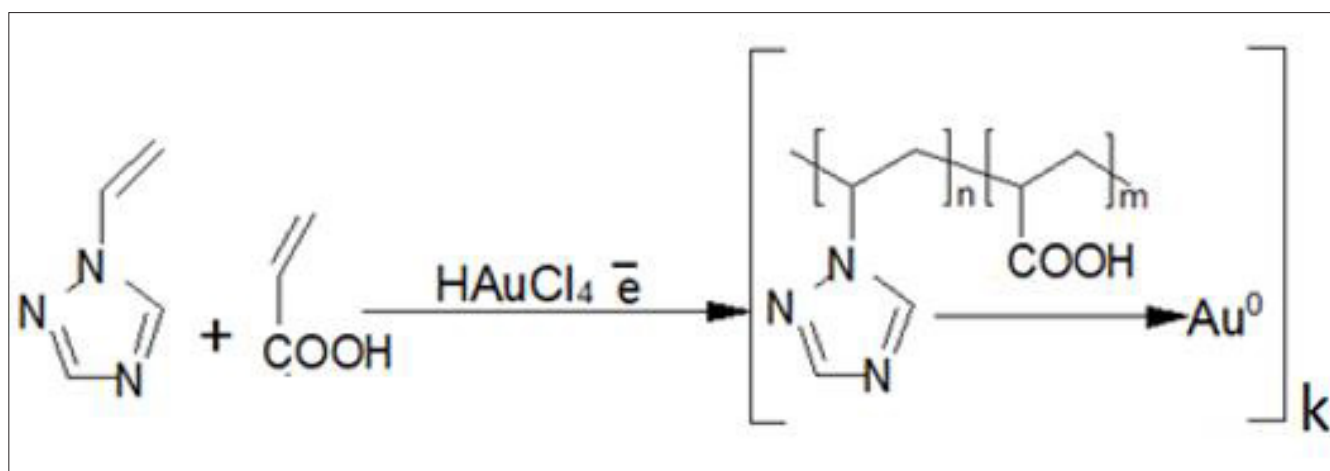


Figure 2: The reaction schematic.

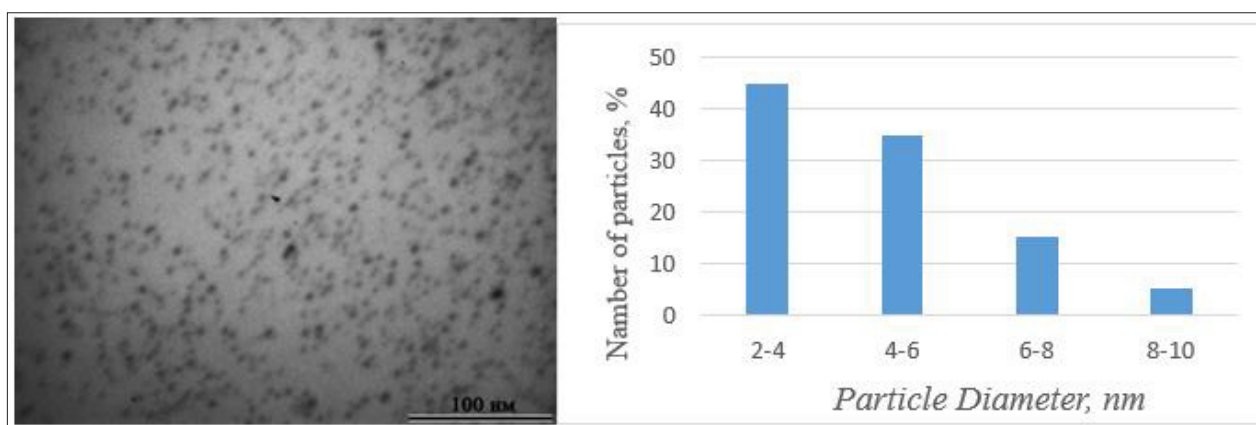


Figure 3: Electron micrograph (a) and distribution diagram (b) gold nanoparticles by size in the VT-AA copolymer matrix.

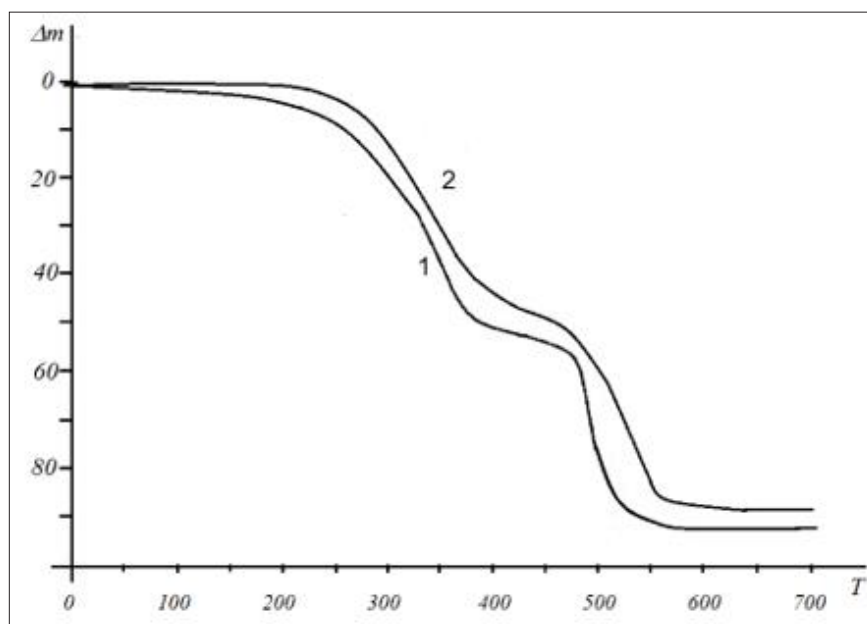


Figure 4: Thermogravimetric curves: 1- copolymer VT-AA, 2 - gold-bearing nanocomposite, - weight loss (mass %), T-temperature (°C).

Conclusion

The possibility of forming nanocomposites and nanocomposite coatings of gold by the electrochemical method based on VT and AA has been shown, and some properties have been studied.

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